****

**CLASS: IX CYCLE TEST- CHEMISTRY MARKS : 20**

**DATE: 9.9.19 Time: 40min**

**I. Answer the following questions:**

1.  The formula for quicklime is\_\_\_\_\_\_\_\_\_ **(1)**

       a.CaCl2 b. CaCo3 c. Ca (OH)2 d. CaO

2. A molecule of oxygen consists of how many atoms of oxygen? **(1)**

a. 6 b.3 c.2 d.0

3. Which of the following statements is true about an atom? **(1)**

a. Atoms are not able to exist independently.

b. Atoms are the basic units from which molecules and ions are formed.



**CLASS: IX CYCLE TEST- CHEMISTRY MARKS : 20**

**DATE: 9.9.19 Time: 40min**

**I. Answer the following questions:**

1.  The formula for quicklime is\_\_\_\_\_\_\_\_\_ **(1)**

       a.CaCl2 b. CaCo3 c. Ca (OH)2 d. CaO

2. A molecule of oxygen consists of how many atoms of oxygen? **(1)**

a. 6 b.3 c.2 d.0

3. Which of the following statements is true about an atom? **(1)**

a. Atoms are not able to exist independently.

b. Atoms are the basic units from which molecules and ions are formed.

c. Atoms are always neutral in nature.

d. Atoms aggregate in large number s to form the matter that we can see, feel or touch.

3. ‘Dalton’s atomic theory is contradicted by the formula of sucrose (C12 H22 O11).’ Justify the statement. **(3)**

4. In a reaction 7.6 g of sodium carbonate reacted with 9 g of ethanoic acid. The products were 3.2 g of carbon dioxide, 1.9 g of water and 11.5 g of sodium ethanoate. Show that these observations are in agreement with the law of conservation of mass. **(3)**

5. Differentiate between atoms and molecules. **(3)**

6. Write the chemical formula for the following: **(3)**

a. Calcium carbonate b. Zinc chloride c. Silver sulphate

7. State the Dalton’s atomic theory. **(5)**

c. Atoms are always neutral in nature.

d. Atoms aggregate in large number s to form the matter that we can see, feel or touch.

3. ‘Dalton’s atomic theory is contradicted by the formula of sucrose (C12 H22 O11).’ Justify the statement. **(3)**

4. In a reaction 7.6 g of sodium carbonate reacted with 9 g of ethanoic acid. The products were 3.2 g of carbon dioxide, 1.9 g of water and 11.5 g of sodium ethanoate. Show that these observations are in agreement with the law of conservation of mass. **(3)**

5. Differentiate between atoms and molecules. **(3)**

6. Write the chemical formula for the following: **(3)**

a. Calcium carbonate b. Zinc chloride c. Silver sulphate

7. State the Dalton’s atomic theory. **(5)**

****

**CLASS: IX CYCLE TEST- CHEMISTRY MARKS : 20**

**DATE: 9.9.19 Time: 40min**

**I. Answer the following questions:**

1.  The formula for quicklime is\_\_\_\_\_\_\_\_\_ **(1)**

       a.CaCl2 b. CaCo3 c. Ca (OH)2 d. CaO

2. A molecule of oxygen consists of how many atoms of oxygen? **(1)**

a. 6 b.3 c.2 d.0

3. Which of the following statements is true about an atom? **(1)**

a. Atoms are not able to exist independently.

b. Atoms are the basic units from which molecules and ions are formed.



**CLASS: IX CYCLE TEST- CHEMISTRY MARKS : 20**

**DATE: 9.9.19 Time: 40min**

**I. Answer the following questions:**

1.  The formula for quicklime is\_\_\_\_\_\_\_\_\_ **(1)**

       a.CaCl2 b. CaCo3 c. Ca (OH)2 d. CaO

2. A molecule of oxygen consists of how many atoms of oxygen? **(1)**

a. 6 b.3 c.2 d.0

3. Which of the following statements is true about an atom? **(1)**

a. Atoms are not able to exist independently.

b. Atoms are the basic units from which molecules and ions are formed.

c. Atoms are always neutral in nature.

d. Atoms aggregate in large number s to form the matter that we can see, feel or touch.

3. ‘Dalton’s atomic theory is contradicted by the formula of sucrose (C12 H22 O11).’ Justify the statement. **(3)**

4. In a reaction 7.6 g of sodium carbonate reacted with 9 g of ethanoic acid. The products were 3.2 g of carbon dioxide, 1.9 g of water and 11.5 g of sodium ethanoate. Show that these observations are in agreement with the law of conservation of mass. **(3)**

5. Differentiate between atoms and molecules. **(3)**

6. Write the chemical formula for the following: **(3)**

a. Calcium carbonate b. Zinc chloride c. Silver sulphate

7. State the Dalton’s atomic theory. **(5)**

c. Atoms are always neutral in nature.

d. Atoms aggregate in large number s to form the matter that we can see, feel or touch.

3. ‘Dalton’s atomic theory is contradicted by the formula of sucrose (C12 H22 O11).’ Justify the statement. **(3)**

4. In a reaction 7.6 g of sodium carbonate reacted with 9 g of ethanoic acid. The products were 3.2 g of carbon dioxide, 1.9 g of water and 11.5 g of sodium ethanoate. Show that these observations are in agreement with the law of conservation of mass. **(3)**

5. Differentiate between atoms and molecules. **(3)**

6. Write the chemical formula for the following: **(3)**

a. Calcium carbonate b. Zinc chloride c. Silver sulphate

7. State the Dalton’s atomic theory. **(5)**